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numbers rules. 1. the  
atoms in a pure  
element must equal  
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Formulas & Chemical  
Compounds Chapter 7

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is the final chapter of the first semester. It serves as a summary of all course content discussed so far, as the focus of this chapter is on process skills rather than new concepts or theory.

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Chemistry Chapter 7

Section 3-4

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Vocabulary. STUDY.

PLAY. Formula mass.

The formula unit of a molecule or ion is the sum of the average atomic masses of all atoms represented in its formula. Percentage composition.

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**Answers Section 4**

CHAPTER 7 REVIEW

Chemical Formulas and  
Chemical Compounds

SECTION 1 SHORT

ANSWER Answer the  
following questions in  
the space provided. 1.

c In a Stock system  
name such as iron(III)  
sulfate, the Roman  
numeral tells us (a)  
how many atoms of Fe  
are in one formula unit.

(b) how many sulfate

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ions can be attached to the iron atom. (c) the charge on each Fe ion.

## **7 Chemical Formulas and Chemical Compounds**

CHAPTER 7 REVIEW

Chemical Formulas and Chemical Compounds

SECTION 1 SHORT

ANSWER Answer the following questions in the space provided. 1. c In a Stock system name such as iron (III) sulfate, the Roman

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numeral tells us (a)  
how many atoms of Fe  
are in one formula unit.  
(b) how many sulfate  
ions can be attached to  
the iron atom.

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Chemical Compounds

SECTION 1 SHORT

ANSWER Answer the following questions in the space provided. 1.

c In a Stock system name such as iron (III) sulfate, the Roman numeral tells us (a) how many atoms of Fe are in one formula unit. (b) how many sulfate ions can be attached to the iron atom.

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SECTION 1 Date

CHAPTER 11 REVIEW

Gases Class SHORT

ANSWER Answer the following questions in the space provided. b Pressure — orce For a constant force, when the surface area is tripled the surface area pressure is (a) doubled. as much. (c ripled. 7-0 (d) unchanged. Rank the following pressures in increasing order. (c) 76 torr (a) 50 kPa O ...

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Name Section Quiz,  
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5. In some instances, the concentration of a solution is expressed as molality instead of molarity because a. molality is easier to calculate. b. molarity applies only to solid-liquid solutions.

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Force. Pressure (P) is defined as the force per unit of area on a surface. pressure = force ÷ area  $P = F/A$ . Atmospheric pressure (atm) is the pressure exerted on an object due to the weight of the column of the air above it in the atmosphere. A .

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1.7 Chapter

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