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Answer: During electrolysis of aqueous

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copper sulphate using copper electrodes, the two anions OH^- and SO_4^{2-} migrate to the anode, but none of them get discharged because the copper of the anode dissolves in the solution producing copper ions and electrons.

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electrodes, of molten compounds (including Lead (II) bromide) and Aqueous solutions (including Sodium Chloride, Dilute Sulfuric Acid and Copper (II) Sulfate) and to predict the products

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The electrolysis of an aqueous solution

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of sodium chloride is shown in Figure. The Na^+ ions and H^+ ions move towards the cathode. The Cl^- ions and OH^- ions move towards the anode. Only one type of ion will be selected to be discharged at the anode and cathode respectively.

Analysing the Electrolysis of

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