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9.1 Reactions and Equations MAIN Idea
Chemical reactions are represented by

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balanced chemical equations. 9.2
Classifying Chemical Reactions MAIN
Idea There are four types of chemical
reactions: synthesis, combustion,
decomposition, and replacement
reactions. 9.3 Reactions in Aqueous
Solutions MAIN Idea Double-replacement
reactions occur between substances

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Chapter 9: Chemical Reactions

CHAPTER Date STUDY GUIDE Class

Chemical Reactions Section 9.1

Reactions and Equations In your textbook, read about evidence of chemical reactions. For each statement, write yes if evidence of a chemical reaction is present. Write no if there is no evidence of a chemical reaction. 2. 4.

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5. 6. 8. A tomato smells rotten.

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Chapter 9 141 Section 9.1 Reactions and
Equations pages 282–288 Practice

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Problems pages 284–287 Write skeleton equations for the following word equations. 1. Hydrogen and bromine gases react to yield hydrogen bromide.

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chemical reaction is present if a tomato
smells rotten t T/F: A chemical reaction
is present if a drinking glass breaks into

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if the percentage yield for the reaction represented by the following equation is calculated to be 75.3%, what mass of Al is expected from the reaction of 52.5g of Al_2O_3

$$2\text{Al}_2\text{O}_3(\text{l}) \rightarrow 4\text{Al}(\text{s}) + 3\text{O}_2(\text{g})$$

Aluminum=26.98 g/mol Oxygen=16.00 g/mol

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(a) A straight horizontal line at 1.0; (b) When real gases are at low pressures and high temperatures, they behave close enough to ideal gases that they are approximated as such; however, in

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some cases, we see that at a high pressure and temperature, the ideal gas approximation breaks down and is significantly different from the pressure calculated by the ideal gas equation.

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Chemistry Matter And Change **Chapter 9 Assessment Answers**

166 Chapter 9 | Stoichiometry Chemical Reactions An example of a chemical reaction is the burning of hydrogen gas (H_2) to produce steam (hot H_2O gas). In chemistry, to burn something is to react it with oxygen gas (O_2) to form

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one or more new substances. Chemical equations are the language used to describe chemical reactions.

Chemistry Chapter 9 Chemical Reactions Answers

Chapter 9 Reactions in Aqueous Solutions • MHR | 2 Net ionic equation:
 $3\text{Ba}^{2+}(\text{aq}) + 2\text{PO}_4^{3-}(\text{aq}) \rightarrow \text{Ba}_3(\text{PO}_4)_2(\text{s})$

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(PO. 4) 2 (s) Check Your Solution . The net ionic equation is balanced, including the charges on the ions. 2. Practice Problem (page 410)

Reactions in Aqueous Solutions

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chemical equation for each of the
follow...

**Solved: What is the general
chemical equation for each of ...**

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The fuel for the space shuttle is

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hydrogen, which burns in oxygen to produce water vapor and energy. In this chemical reaction, hydrogen is a(n) (9), oxygen is a(n) (10), and water vapor is a(n) (11). In a chemical equation for this reaction, a(n) (12) is used to separate hydrogen and oxygen from water vapor and energy.

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Engage Discuss the temperature changes in chemical reactions students have conducted so far Chapter 9 chemical reactions answer key. Remind students that the decomposition reaction of hydrogen peroxide and the reaction with copper II sulfate and aluminum both caused the temperature of the solution to increase. Chapter 9

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Summary of the Reactions of Chapter 9:

1) Reactions of Carboxylic Acids a)

Reaction of a Carboxylic acid and water.

- When placed in water, a carboxylic acid molecule acts as an acid and water

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acts as a base.

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