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Section 11.1 What is stoichiometry? 1.

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## Answers

true 2. true 3. false 4. true 5. true 6. 2,  
2, 64.10 7. 3, 3, 96.00 8. 2, 2, 88.02 9. 4,  
4, 72.08 10. methanol and oxygen gas  
11. carbon dioxide and water 12. 160.10  
g 13. 160.10 g 14. They are equal. 15. A  
mole ratio is a ratio between the  
numbers of moles

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mole ratio Concepts Using Balanced  
Equations • Number of Atoms • Number  
of Molecules • Number of Moles • Mass o  
Law of Conservation of Mass • Volume

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11.2 Stoichiometric Calculations  
Concepts

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368 Chapter 11 • Stoichiometry Section  
11.11.1 Objectives Describe the types  
of relationships indicated by a balanced  
chemical equation. State the mole ratios

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from a balanced chemical equation.

Review Vocabulary reactant: the starting substance in a chemical reaction

New Vocabulary stoichiometry mole ratio

Defining Stoichiometry

## **Chapter 11: Stoichiometry**

Stoichiometry The study of quantitative relationships between the amounts of



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reactants used and amounts of products formed by a chemical reaction is called stoichiometry. Stoichiometry is based on the law of conservation of mass. Recall that the law states that matter is neither created nor destroyed in a chemical reaction.

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Chapter 11 Stoichiometry.

stoichiometry. mole ratio. excess reactant. limiting reactant. The study of quantitative relationships between the amounts of.... In a balanced equation, the ratio between the numbers of moles.... A reactant that remains after a chemical reaction stops.

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Equations • Number of Atoms • Number of Molecules • Number of Moles • Mass of Law of Conservation of Mass • Volume  
11.2 Stoichiometric Calculations  
Concepts Mole-Mole...

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Change • Chapter 11 209  
Stoichiometry Stoichiometry CHAPTER 11

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## Answers

SOLUTIONS MANUAL Section 11.1  
Defining Stoichiometry pages 368–372  
Practice Problems pages 371–372 1.  
Interpret the following balanced  
chemical equations in terms of  
particles, moles, and mass. Show that  
the law of conservation of mass is

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...

The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry. Stoichiometry is based on the law of conservation of mass, In any chemical reaction, the



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mass of the products is equal to the mass of the reactants.

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Stoichiometry Chapter 11 Study Guide  
Answer Key Stoichiometry is the tool for  
answering these questions.  
Stoichiometry The study of quantitative

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relationships between the amounts of reactants used and amounts of products formed by a chemical reaction is called stoichiometry. Stoichiometry is based on the law of conservation of mass.

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CHAPTER Section 11.1 continued In your

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textbook, read about mole ratios.

Answer the questions about the following chemical reaction. sodium + iron(III) oxide  $\rightarrow$  sodium oxide + iron

$6\text{Na}(s) + \text{Fe}_2\text{O}_3(s) \rightarrow 3\text{Na}_2\text{O}(s) + 2\text{Fe}(s)$

15. What is a mole ratio? 16. How IS a mole ratio written? CA S Q C CYA 17. Predict the number of mole ratios for this reaction. Class 18.

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## Answers

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Section 111 What is stoichiometry? 1  
true 2 true 3 false 4 true 5 true 6 2, 2,  
6410 7 3, 3, 9600 8 2, 2, 8802 9 4, 4,  
7208 10 methanol and oxygen gas 11  
carbon dioxide and water 12

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Stoichiometry. study of quantitative relationships between the amounts of reactants used and the amounts of products formed in a chemical reaction.

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